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Percent Composition Review Answer Key

Percent Composition 2. What is the percent composition of a compound

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that contains 8.1 grams of nickel and 2.2 grams of oxygen in a 10.3-gram sample? Nickel:  $100 \frac{8.1}{10.3} = 78.6\%$  Oxygen:  $100 \frac{2.2}{10.3} = 21.4\%$  3. Nicotine, the addictive drug in cigarettes, contains 74.0% carbon, 8.6% hydrogen, and 17.3% nitrogen.

Instructors Guide: Percent Composition

Percent Composition by Mass

Worksheet Experimental Percent

Composition by Mass 1. A chemist

determines that 1.26 g of iron reacts with 0.54 g of oxygen to form rust.

What is the percent composition of each element in the new compound?

$$\text{Fe}\% = \left(\frac{m \text{ Fe}}{m \text{ total}}\right) * 100$$

$$\left(\frac{1.26}{1.26+0.54}\right) * 100 = 70.0\%$$

$$\text{O}\% = \left(\frac{m \text{ O}}{m \text{ total}}\right) * 100$$

$$\left(\frac{0.54}{1.26+0.54}\right) * 100 = 30.0\%$$

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Percent Composition by Mass Worksheet Experimental Percent Composition by Mass 1. A chemist determines that 1.26 g of iron reacts with 0.54 g of oxygen to form rust. What is the percent composition of each element in the new compound?

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$$\text{Fe}\% = \left( \frac{m \text{ Fe}}{m \text{ total}} \right) * 100$$
$$(1.26 / (1.26 + 0.54)) * 100 =$$

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( percent composition of the elements in the compound (3) heat of fusion of the compound X (4) melting point of the elements in the compound\* prices from the two suppliers are the same for a 500.0-g bottle, which one would supply more of the desired Na<sub>2</sub>B<sub>4</sub>O<sub>7</sub>? Explain your answer. Title: new doc 2019-02-13 10.43.55

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Percent Composition Review Answer Key  
composition is as follows, what is the empirical and molecular formula of

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serine? C = 34.95 % H= 6.844 % O = 46.56 % N= 13.59 % Percent Composition and Molecular Formula Worksheet Key

Percent Composition and Molecular Formula Worksheet

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Percent Composition Worksheet Answer Key with Work ...

Percent Composition of a Compound  
Percent composition is the percent by mass of each element in a compound. To find the percent by mass of an

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Answer Key  
element in a compound, use the formula:  $\frac{\text{mass of element}}{\text{mass of compound}} \times 100\%$

## 10.3 Percent Composition and Chemical Formulas

To calculate percent composition, we divide the experimentally derived mass of each element by the overall mass of the compound, and then convert to a percentage:

$$\left[ \frac{\text{mass of element}}{\text{mass of compound}} \times 100\% \right]$$

Example:  $\frac{7.34 \text{ g C}}{12.04 \text{ g compound}} \times 100\% = 61.0\%$

## 4.2: Formula Mass, Percent Composition, and the Mole ...

For these sorts of applications, the percent composition of a compound is easily derived from its formula mass and the atomic masses of its

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constituent elements. A molecule of  $\text{NH}_3$  contains one N atom weighing 14.01 amu and three H atoms weighing a total of  $(3 \times 1.008 \text{ amu}) = 3.024 \text{ amu}$ .

5.4 Percent Composition, Empirical and Molecular Formulas ...

EMPIRICAL/MOLECULAR  
FORMULA/PERCENT

COMPOSITION 5. Find the empirical formula of a compound which is 49.5% Fe, and 50.5% F. □ Step 1: Convert to grams Fe: 49.5 grams F: 50.5 grams  
Step 2: Convert grams to Moles Fe: 49.5 grams = 0.887 M F: 50.5 grams = 2.66 M 55.8 grams 19.0 grams  
Step 3: Divide by Smallest Moles

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Here's a great practice worksheet that includes 10 questions. Part A of this worksheet gives your students practice with finding the percent composition of a compound from a given formula, finding percent composition of a compound from a chemical name, and finding percent water in a hydrate.

Percent Composition, Empirical Formulas, & Molecular ...

Percent Composition Name Date  
Show all work for finding the molar mass AND the percent composition requested. Example: What percent

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of NaCl is sodium? Answer: Molar mass is  $22.99 + 35.45 = 58.44$  g/mol  
Percent composition: %Na =  $(22.99 / 58.44) \times 100\% = 39.34\%$  % Na  
1) What percent of  $\text{MgB}_2$  is magnesium? Page 1 of 1  
Hour 0) Molar Mass =

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