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Types of Chemical Reactions

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STUDY GUIDE Section 9.2 Classifying
Chemical Reactions In your textbook,
read about synthesis, combustion,
decomposition, and replacement
reactions. Assume that Q, T, X, and Z
are symbols for elements. Match each
equation in Column A with the reaction

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type it represents in Column B.

Column A 2. Q + Z → QZ Answer the following questions. a. b. c. d.

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In a chemical reaction, chemical (blank) are formed or broken exothermic; product During an (blank) reaction, more energy is required to form the bonds in the (blank) than needed to break the bonds of the product are formed.

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Concentration increase, less surface area, higher temperature, and a catalyst. Four ways a chemical reaction has taken place. Color change, the formation of a precipitate, gas formed, temperature change. 3

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examples of a physical change.

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replaces a second element in a

compound. question. double

replacement reaction. answer. A

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exchange of positive ions between two compounds. question. combustion reaction.

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o u u o o 0 o o 0 o o o 0 2 0 o N x N u

N o m g g 0 0 o O o N O N O m o z N 9

o r r N r z

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Textbook. Used to show what a

chemical reaction yields (same as =);

used to separate the reactants from

the product. Left side items are

reactants and right side items are the

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product. (g) = gas; (aq) = water; (l) = liquid. these symbols used to denote the state of the reactants and products in an accurate balanced chemical equation.

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1. State the term for the amount of energy that is needed for a chemical reaction to occur. 2. Summarize the relationship between an enzyme and a substrate. enc 3. Classify which of the compounds in the reaction below are reactants and which are products.

$PbO_2 + 4HCl \rightarrow PbCl_2 + 2H_2O$ 4.

Compare endothermic reactions and exothermic reactions. f 5.

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Chapter 10 Study Guide Chemical Reactions Answer Key

View Test Prep - Chemical reactions study guide from CHEM 1100 at University Of Georgia. Unit 6 Study Guide for Chemical Reactions Please answer the following questions: 1. What is a chemical

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: Answer the questions and complete statements from viewing the video. Everything is made of chemicals. When do chemical reactions happen? When the electrons hook together. Water is . two. parts H and . one. part O . What makes iron rust and our blood red? Oxygen . Car and motorcycle engines are controlled by what chemical reaction?

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Topics you will need to know in order to pass the quiz include reactants and chemical ...

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why the formation of

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For most chemical reactions in this book, solids are labeled (s), liquids (l), and gases (g). The numerical coefficients in front of the chemical formulas express the moles of each compound or element. The preceding reaction can be interpreted in terms of moles or masses. (See Table 1.)

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When a chemical reaction occurs, the original substances put together, called reactants, lose their chemical properties and become different substances called products with a different set of chemical properties. Reactions where energy is released

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are called exothermic reactions. When energy is absorbed, it is called an endothermic reaction.

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Ritter, Plumb, et al; Nelson 2001.

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memorization, and true understanding comes only from concentrated reading, and working as many problems as possible. In fact, problem sets are the best way to ensure that concepts are not only well understood, but can also be applied to real-world problems in the work place. Helps readers learn to categorize, analyze, and solve organic chemistry problems at all levels of difficulty Hundreds of fully-worked practice problems, all with solutions Key concept summaries for every chapter reinforces core content from the companion book

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