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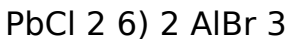
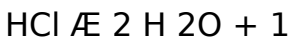
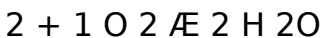
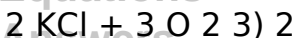
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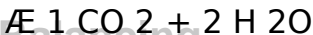
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below: 1) $\underline{\quad}$ N_2 +
 $\underline{\quad}$ $\text{H}_2 \rightleftharpoons \underline{\quad}$ NH_3

... 7) Copper and
sulfuric acid react
to form copper (II)
sulfate and water
and sulfur dioxide.

8) Hydrogen gas
and nitrogen
monoxide react to

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form water and
nitrogen gas.

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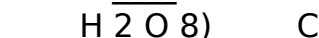
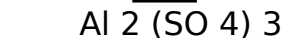
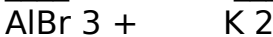
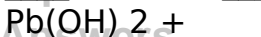
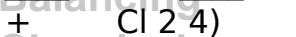
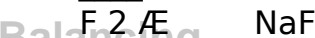
below: 1) $\underline{\quad}$ N₂

+ $\underline{\quad}$ H₂ \rightarrow $\underline{\quad}$

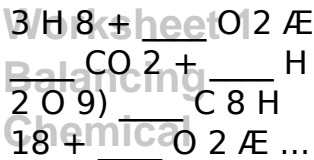
NH₃ 2) $\underline{\quad}$ KClO₃

\rightarrow $\underline{\quad}$ KCl + $\underline{\quad}$ O₂

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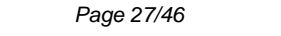
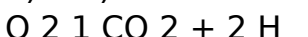
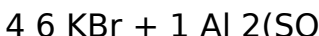
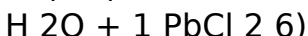
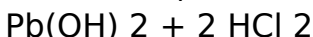
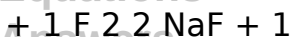
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below: 1) $_ \text{N}_2 +$

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Key Balance the
equations below: 1)



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20 8) $1 \text{ C}_3\text{H}_8 + 5 \text{ O}_2 \rightarrow 3 \text{ CO}_2 + 4 \text{ H}_2\text{O}$
20 9) $2 \text{ C}_8\text{H}_{18} + 25 \text{ O}_2 \rightarrow 16 \text{ CO}_2 + 18 \text{ H}_2\text{O}$
20 10) $1 \text{ FeCl}_3 + 3 \dots$

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change the value of coefficient in front of the element or compound, and not the subscript.

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product sides. This is a requirement the equation must satisfy to be consistent with the law of

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have the strength
to focus on these
tasks.

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absorbs energy
from its

surroundings. Law

of Conservation of

Energy energy

cannot be created

or destroyed, but it

can be transferred

or transformed

from one form to

another. Chapter 7

Balancing Chemical

equation

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As illustrated in
Figure 7.3.2, the
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us to interpret

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Equation 7.3.1 in any of the following ways: Two NH_4^+ ions and one $\text{Cr}_2\text{O}_7^{2-}$ ion yield 1 formula unit of Cr_2O_3 , 1 N_2 molecule, and 4 H_2O molecules. One mole of $(\text{NH}_4)_2\text{Cr}_2\text{O}_7$ yields 1 mol of Cr_2O_3 , 1 mol of N_2 , and 4 mol of H_2O .

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