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Periodic Table Trends
– Electron Configs
and Ionization

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Properties of the
Elements: Part 3 of 11
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unreactive due to
electron configuration~~

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- ns^2np^6 (except He $1s^2$) – Main group elements tend to gain or lose electrons to become isoelectronic (same valence electron configuration as nearest noble gas)

~~Chapter 7 Electron
Configuration and the
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- 5 - between the electron in that orbital and the nucleus. This is mainly a function of the average distance of the electron from the nucleus, which is controlled by the principal quantum number, n .

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~~Chapter 7~~
the Properties of
Atoms

18 7.4 Periodic
Trends in Properties
of Elements • Atomic

radius: distance
between nucleus of
an atom and its
valence shell •

Metallic radius: half
the distance between
nuclei of two adjacent,
identical metal atoms

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~~Chapter 7 Electron Configuration and the Periodic Table~~

As you have learned,
the electron

configurations of the
elements explain the
otherwise peculiar
shape of the periodic
table. Although the
table was originally
organized on the
basis of physical and
chemical similarities

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between the elements within groups, these similarities are ultimately attributable to orbital energy levels and the Pauli principle, which cause the individual subshells to be ...

~~7.8B: Electron
Configurations and
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7 Electron
Configurations
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Predicting Chemical
Properties Valence
electrons are the
outermost electrons
and are involved in
bonding. Chapter 7
Section 2 Similarity of
valence electron
configurations helps
predict chemical

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Chapter 5
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properties. Groups
1A, 2A and 8A all
have similar
properties to other
members of their
respective groups.

~~Electron Configuration and the Periodic Table~~

The actual electron
configuration may be
rationalized in terms
of an added stability

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associated with a half-filled ($ns\ 1$, $np\ 3$, $nd\ 5$, $nf\ 7$) or filled ($ns\ 2$, $np\ 6$, $nd\ 10$, $nf\ 14$)

subshell. (In fact, this "special stability" is really another

consequence of the instability caused by pairing an electron with another in the same orbital, as ...

~~7.7B: Electron~~

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Electron

Configurations for

Anions • Nonmetals

gain electrons to
attain noble gas

configuration. • They
make negative ions. •

If we look at electron
configuration it makes
sense. • S $1s^2 2s^2$

$2p^6 3s^2 3p^4$ 6

valence electrons •

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S-2 1s 2 2s 2 2p 6 3s
2 3p 6 [Ar] noble gas
configuration.

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Explanations of
periodic trends in
physical and chemical
properties are based

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Chapter 7

configurations –
arrangements of
electrons within atoms

7.1 Electromagnetic Radiation and Matter

- Excited – atoms gain energy when electrons absorb light and then release it in form of electromagnetic radiation (some is visible) •

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Electromagnetic radiation consists of oscillating perpendicular electric and magnetic fields that travel ...

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Electron Configurations and the Properties of Atom

The electron configurations described in this chapter all refer to gaseous atoms in their ground states. An atom may absorb a quantum of energy and promote one of

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its electrons to a higher-energy orbital. When this happens, we say that the atom is in an excited state. The electron configurations of some excited atoms are given.

~~Answer: The electron configurations described in this ...~~

Determine which ions

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have noble-gas configurations. Check all that apply. Zr^{4+} S^{2-} ? Part A Select the atoms whose ions have an electron configuration of nd^6 ($n = 3, 4, 5 \dots$). For example, the $2+$ ion for V ends the electron configuration with $3d^3$. Check all that apply. Co Pt Ru.
Part A

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~~Chapter 5— Electron
Configurations and
the Periodic Table ...~~

Write the electron
configuration for zinc.
15. Fill in the electron
configuration diagram
for the copper(I) ion.

Copper atom Cu 3d
4s 3p 3s 2p 2s 1s

Energy level

Copper(I) ion Cu

Formation of Anions

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16. Atoms of most nonmetallic elements achieve noble-gas electron configurations by gaining electrons to become , or negatively charged ions. 17.

~~BONDING AND INTERACTIONS~~

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~~Alkaline Earth Metals
Electron Configuration
The Earth ...~~

The periodic table can be a powerful tool in predicting the electron configuration of an element. However, we do find exceptions to the order of filling of orbitals that are shown in Figure 3 or Figure 4. For instance,

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the electron configurations (shown in Figure 6) of the transition metals

chromium (Cr; atomic number 24) and

copper (Cu; atomic number 29), among

others, are not those we would ...

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