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Acids Bases
Chapter 15

Answers
1 Acids

Bases

Answers

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well as
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lesson,
amusement, as
well as harmony

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to Acids and
Bases *Organic
Chemistry Acids*

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*Acids Bases –
Reactions,
Strength,
Acidity, Pka*
\u0026

Conjugates ACIDS
BASES \u0026
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~~Chapter 16 (Acid-Base Equilibria)~~

~~Part 1 Chapter~~

15 - Acid Base

Equilibria -

Part I Chapter

16 - Acid-Base

Equilibria: Part

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~~IMPORTANT~~

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~~CHAPTER 2 ACIDS,~~

~~BASES AND SALTS~~

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Chapter 16 Acid-

Base Equilibria

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General

Properties of

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Chemistry Acid

Base and Salts

// Chapter 2 //

Reactions of

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Acids | | Class

10 | | Part 4

Answers

Acids and Bases
and Salts -

Introduction |

Chemistry |

Don't Memorise

Easy way to
memorize the 7
strong acids and
6 strong bases

Acids and Bases,

pH and pOH ~~Acids~~

~~and Bases~~

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~~Chemistry~~

~~Basic~~

~~Introduction~~

~~Acid-Base~~

~~Equilibria and~~

~~Buffer Solutions~~

~~Acid-Base~~

~~Titration~~

~~Curves, pH~~

~~Calculations,~~

~~Weak \u0026amp;~~

~~Strong,~~

~~Equivalence~~

~~Point, Chemistry~~

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~~Problems~~

~~CHY 115: Acid-
Base Equilibrium~~

~~Calculation~~

~~Problems~~

~~Lewis~~

~~Concept About~~

~~Acids \u0026~~

~~Bases, Acids,~~

~~Bases, and pH~~

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~~Chapter 15 Acids~~

~~and Bases, A~~

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2 Chemistry 102:
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and Bases, A
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Chapter 14
(Acids and
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~~Acids Bases and
Salts 10th Class
Chemistry,~~

~~Bronsted Lowry~~

~~Concept - Ch 10~~

~~-Matric Class~~

~~Chemistry Acids,~~

~~Bases \u0026~~

~~Salts (?????)~~

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~~10 Science Acid~~

~~Base Chapter 15~~

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Salts - ep03 -

BKP | class 10

Answers
science ch 2

explanation in

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chemistry

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~~Acids Bases~~

15.1 Acids and

bases acid

derived from

Latin acidus

(meaning sour or

tart) related to

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Latin acetum

(meaning
vinegar)

characteristic
properties
associated with
acid: 1. sour
taste 2. change
the color of
litmus from blue
to red 3. react
with ?metal
(such as Zn, Mg)
to produce H₂

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gas?hydroxide
base to produce
H₂O and salt
?carbonate to
produce CO₂

~~Chapter 15~~

~~Acids, Bases and
Salts~~

Acids and Bases.
GCh15-1. Chapter
15. Acids and
Bases. What we
will learn:

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- Bronsted acids and bases
- Acid-base properties of water
- pH - a measure of acidity
- Strength of acids and bases
- Weak acids (bases) and acid (base) ionization constant
- Diprotic and

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polypropionic acids

- Molecular structure and strength of acids
- Acid-base properties of salts
- Acid-base properties of oxides and hydroxides
- Lewis acids and bases.

~~Chapter 15.~~

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~~Acids and Bases~~

15 Drill:

Answers

Identify the B-L
acid and base in
each of the
following.

Circle any
amphoteric
species acid
acid base base

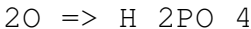
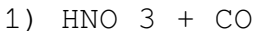
Note: In both
examples, water
behaved as an
acid or a base.

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A species that can act as an acid or a base is called amphoteric.

Bronsted-Lowry (B-L) Theory -
Cont. acid base



~~Chapter 15 Acids~~

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Chapter 15 1

~~and Bases~~

~~Bridgewater~~

~~State University~~

Sec 15.3:

Definitions of

Acids and Bases:

Arrhenius

Definition-

Acids: a

substance that

produces H^+ ions

in an aqueous

solution. Bases:

a substance that

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Chapter 15 1

produces OH^- ions in aqueous solution. *Note: under this definition, the acid and base combine to form water, neutralizing each other in the process

~~Chemistry~~

~~Chapter 15 Acids~~

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~~and Bases Book~~

~~Notes CH 232~~

~~Answers~~

...

Chapter 15

Review Acids

Bases 1Chapter

15 Acids and

Bases Multiple

Choice Questions

1) _____ is

found in

carbonated

beverages due to

the reaction of

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carbon dioxide
with water.

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and Bases -

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and Bases.

strong acid.

strong base.

weak acid. weak

base. an acid

that ionizes

completely ...

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Acids Bases

~~Chapter 15~~

~~Review Acids~~

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ACIDS AND BASES

453 SECTION 15-1

OBJECTIVES List

five general
properties of
aqueous acids
and bases. Name
common binary
acids and

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oxyacids, given their chemical formulas. List five acids commonly used in industry and the laboratory, and give two properties of each. Define acid and base according to Arrhenius's theory of

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ionization.

Explain the
differences

~~CHAPTER 15 Acids
and Bases~~

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acids bases

answers, but

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manner of a mug
of coffee in the
afternoon,
instead they
juggled behind
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virus inside
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~~Acids Bases~~

~~Answers~~

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Acids and Bases

Acids and Bases

Arrhenius

De?nitions:

acids -

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Acids - compounds that produce an increase in $[H^+]$ when dissolved in water
Bases - compounds that produce an increase in $[OH^-]$ when dissolved in water
Lewis
Definitions:
acids - electron pair acceptors

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bases – electron

pair donors

Brønsted–Lowry

Definitions:

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~~Acids and Bases~~

~~Acids and Bases~~

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Answers

Recognizing the
showing off ways
to get this

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1 acids bases

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STUDY. PLAY.

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idea. acids and
bases change the
color of
compound
indicators. ...

-Lewis acid-base
reaction is the
formation of one
or more covalent
bonds between an
electron-pair
donor and an

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electron pair

acceptor,

although the 3

acid-base

reactions

differ, many

compounds can be

categorized as

acids ...

~~Chapter 15 Acids~~

~~and Bases~~

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Chapter 15 Acids and Bases

Multiple Choice Questions 1)

_____ is found
in carbonated
beverages due to
the reaction of
carbon dioxide
with water. A)
 CH_3COOH ; B)
 H_2CO_3 ; C) HCOOH ;
D) $\text{C}_6\text{H}_5\text{COOH}$; E)
 $\text{CH}_3\text{CH}_2\text{COOH}$;

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Answer: B. Diff:

1 Page Ref: 15.2

2) Which of the following species is

amphoteric? A)

CO_3^{2-} B) HF; C)

NH_4^+ ? D) HPO_4^{2-}

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feel? Have you
felt right?
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interest at
once.

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Answers Chapter

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on acids and

bases, we saw

two more

definitions of

acids: a

compound that

donates a proton

(a hydrogen ion,

H^+) to another

compound is

called a

Brønsted-Lowry

acid, and a

Lewis acid is

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any species that
can accept a
pair of
electrons.

~~15.2 Lewis Acids
and Bases~~

~~Chemistry~~

Conjugate Acid

Base Pairs,

Arrhenius,

Bronsted Lowry

and Lewis

Definition -

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Chemistry –
Acids Bases

Answers
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