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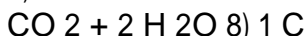
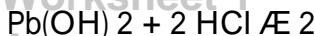
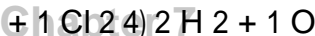
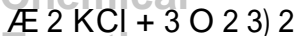
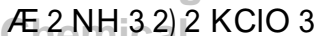
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$18 + 25 \text{ O}_2 \rightarrow 16 \text{ CO}_2 + 18 \text{ H}_2\text{O}$...

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Balancing Equations. When a chemical equation is balanced it means that equal numbers of atoms for each element

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involved... Additional
Information in
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The physical states of
reactants and products
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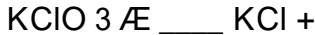
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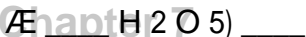
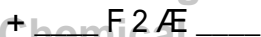
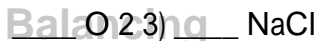
Equations Balance the

equations below: 1)



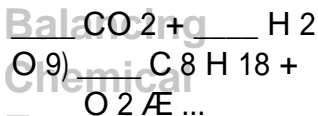
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chemical reaction that
absorbs energy from its
surroundings. Law of
Conservation of Energy
energy cannot be
created or destroyed,
but it can be transferred

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or transformed from one form to another.

Chapter 7 Balancing

Chemical equation

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Balancing Equations. A

Answers
balanced chemical is

equation has equal

numbers of atoms for

each element involved

in the reaction are

represented on the

reactant and product

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sides. This is a requirement the equation must satisfy to be consistent with the law of conservation of matter. It may be confirmed by simply summing the numbers of atoms on either side of the arrow and comparing these sums to ...

7.1: Writing and

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Balancing Chemical Equations | General ...
Answer Key Chapter 7 -
Chemistry: Atoms First
2e | OpenStax. 1. An equation is balanced when the same number of each element is represented on the reactant and product sides. Equations must be balanced to accurately reflect the law of conservation of matter.

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3. Balancing

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Balancing chemical equations involves the addition of

stoichiometric coefficients to the reactants and products.

This is important because a chemical equation must obey the

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law of conservation of mass and the law of constant proportions , i.e. the same number of atoms of each element must exist on the reactant side and the product side of the equation.

Balancing Chemical Equations - BYJUS

The first step to balance the equation is to write

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down the chemical formula of reactants that are listed on the left side of the chemical equation. After this, you can list down the products on the right hand side of the chemical equation. There is an arrow between the sides, signaling the direction the reaction is happening in.

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Chapter 7.1-7.2

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Balancing Equations.

The chemical equation
described in section 4.1
is balanced, meaning
that equal numbers of

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atoms for each element involved in the reaction are represented on the reactant and product sides. This is a requirement the equation must satisfy to be consistent with the law of conservation of matter.

7.1 Writing and Balancing Chemical Equations – General ...

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Check the chemical
equation to make sure it
is balanced as written;
balance if necessary.

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Then calculate the number of moles of $[\text{Au}(\text{CN})_2]^-$ present by multiplying the volume of the solution by its concentration. From the balanced chemical equation, use a mole ratio to calculate the number of moles of gold that can be obtained from the reaction.

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equation has equal numbers of atoms for each element involved in the reaction are represented on the reactant and product sides. This is a requirement the equation must satisfy to be consistent with the law of conservation of matter. It may be confirmed by simply summing the numbers

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of atoms on either side
of the arrow and
comparing these sums to

...

Chapter 7

7.2 The Chemical
Equation: Balancing
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When balancing
equations, remember
chemical reactions must
satisfy conservation of
mass. Check your work
to make certain you

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have the same number and type of atoms on the reactants side as on the products side. A coefficient (number in front of a chemical) is multiplied by all the atoms in that chemical.

Balancing Equations

Chemistry Test

Questions

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Equations ...

Balancing Chemical
Equations – Answer
Key Balance the

equations below: 1) $1 \text{ N}_2 + 3 \text{ H}_2 \rightarrow 2 \text{ NH}_3$ 2) $2 \text{ KClO}_3 \rightarrow 2 \text{ KCl} + 3 \text{ O}_2$
3) $2 \text{ NaCl} + 1 \text{ F}_2 \rightarrow 2 \text{ NaF} + 1 \text{ Cl}_2$ 4) $2 \text{ H}_2 + 1 \text{ O}_2 \rightarrow 2 \text{ H}_2\text{O}$ 5) 1 Pb(OH)_2

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+ 2 HCl 2 H₂O + 1
PbCl₂ 6) 2 AlBr₃ + 3
K₂SO₄ 6 KBr + 1 Al
2(SO₄)₃ 7) 1 CH₄ + 2
O₂ 1 CO₂ + 2 H₂O
8) 1 C₃H₈ + 5 O₂ 3
CO₂ + 4 H₂O 9) 2 C
8H₁₈ + 25 O₂ 16 CO
2 + 18 H₂O 10) 1 FeCl
3 + 3 ...

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